

Yangi mavzuning bayoni

Reja:

1. Oksidlanish qaytarilish reaksiyalari.
2. Oksidlanish darajasi.
3. Oksidlanish qaytarilish reaksiyalari tenglamalarini tuzish usullari.

Ushbu reja bilan talabalar oldingi darsda tanishgan. Yangi darsda reja asosida tayyorlanid kelish talabalarga topshiriq sifatida tayinlangan. Shuning uchun ularni ikki guruhga bo'lib o'yin viktorina tashkil qilinadi. Har bir to'g'ri javob bergan komandaga bir baldan berib boriladi. Kim ko'p javob besa shu o'quvchi komanda sardori etib tayinlanadi. Bunda sardorlar bellashuvi boshlanadi.

Viktorina savollari:

I-komandaga beriladigan savollar:

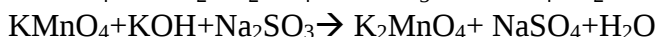
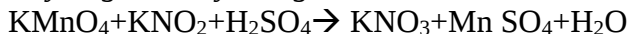
1. Ham oksidlovchi, ham qaytaruvchi bo'la oladigan elementlarga misol keltiring.
2. KMnO_4 neytral muhitda qanday oksidlanish darajasini hosil qiladi?
3. Oksidlanish darajasiga ta'rif bering.
4. Ion-elektron usuli deganda nimani tushunasiz va uning afzalligi nima?
5. Oksidlovchilarga misol keltiring va uning oksidlovchlik xossalari saababini tushuntiring.
6. Oksidlanish qaytarilish reaksiyalarining mohiyati nima?
7. Oksidlovchi va qaytaruvchi nima?

II- komandaga beriladigan savollar:

1. D.I.Mendeleevning elementlar davriy sistemasiga qarab qaysi elementlar oksidlovchi va qaysilari qaytaruvchi bulishini tushuntiring.
2. KMnO_4 ishqoriy muhitda qanday oksidlanish darajasini hosil qiladi?
3. Oksidlanish qaytarilish reaksiyalarining qanday turlari bor.
4. Oksidlanish qaytarilish reaksiyalarining kundalik hayotda tutgat o'rnini izohlang.
5. Ftordan yodga o'tgan sari galogenlarning oksidlovchlik xossasi qanday o'zgaradi?
6. Oksidlanish qaytarilish reaksiyalariga muhitning qanday ta'siri bor? Misollar bilan tushuntiring.
7. Metallar va metalmaslarga umumiy tarif bering.

Komanda sardorlariga topshiriq:

Quyidagi reakdiya tenglamalarini ion-elektron usuli bilan tenglashtiring.



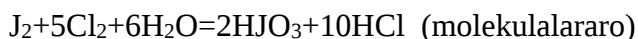
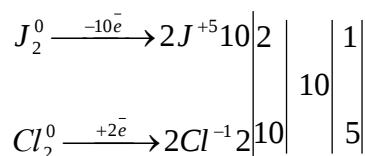
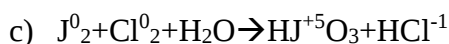
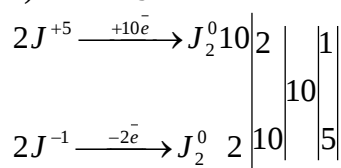
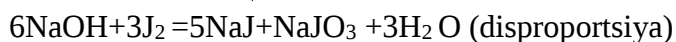
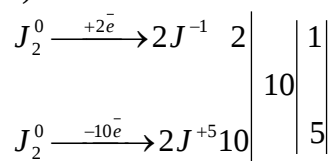
(20 minut)

Yangi mavzuni mustahkamlash:

Masala va mashqlar echish.

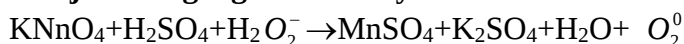
Quyidagi reaksiyalar oksidlanish – qaytarilish reaksiyalarining qaysi turiga kiradi? Reaksiyani electron balans usulida tenglashtirilgan tenglamasini tuzing.

- a) $\text{NaOH} + \text{J}_2 \rightarrow \text{NaJ} + \text{NaJO}_3 + \text{H}_2\text{O}$
- b) $\text{KJO}_3 + \text{KJ} + \text{H}_2\text{SO}_4 \rightarrow \text{J}_2 + \text{K}_2\text{SO}_4 + \text{H}_2\text{O}$
- c) $\text{J}_2 + \text{Cl}_2 + \text{H}_2\text{O} \rightarrow \text{HJO}_3 + \text{HCl}$

Yechimi:

(30 minut)

Uy vazifa: O'tilgan darsda bajarilgan tajribalarga tegishlin reaksiyalar tenglamasini tuzish va elektron balans usulida tenglashtirish. 3 ta o'quvchi uy vazifasini doskada bajaradi. Qolganlarining daftari tekshirib chiqiladi.

1-tajribaning tegishli reaksiyasi:**2- tajribaning tegishli reaksiyasi:****3- tajribaning tegishli reaksiyasi:****Adabiyotlar ro'yxati:**

1. M.M Abdulxayeva, O'.M. Mardonov – Kimyo T"O'z" 2003yil
2. I.Tirkashov, S. Masharipov – Kimyo T "O'qituvchi"2003 yil
3. G.P. Xomchenko – Kimyo Oliy o'quv yurtiga kiruvchilar uchun T "O'q"2001 yil

(5 minut)

Navbatdagi mavzu rejasi bilan tanishish**Reja:**

1. Elektroliz tushunchasi bilan tanishish
2. Faradey qonunlari
3. Elektrolizning amaliy ahamiyati

(3 minut)